

Molecular Biology Grade 1M Tris-HCl Buffer

| Product | Con. | Cat# | Size |
|-----------------|------------|-------------|-------|
| Tris-HCl Buffer | 1M, pH 7.2 | IBS-BT016 | 500ml |
| | 1M, pH 7.0 | IBS-BT016-1 | 500ml |
| | 1M, pH 7.4 | IBS-BT017 | 500ml |
| | 1M, pH 7.4 | IBS-BT017b | 1L |
| | 1M, pH 7.6 | IBS-BT018 | 500ml |
| | 1M, pH 8.0 | IBS-BT019 | 500ml |
| | 1M, pH 8.3 | IBS-BT020 | 500ml |
| | 1M, pH 8.3 | IBS-BT020-1 | 1L |
| | 1M, pH 8.5 | IBS-BT021 | 500ml |
| | 1M, pH 9.0 | IBS-BT022 | 500ml |
| | 1M, pH 8.8 | IBS-BT080 | 500ml |

Components : 1M Tris (hydroxymethyl) aminomethane buffer, pH x.x
pH adjusted with hydrochloric acid(HCl), sterile solution

Storage Conditions : Room Temperature
Stable for a minimum of 1 year from date of receipt at room temperature.

Introduction : Tris (also known as THAM) is an abbreviation of the organic compound known as tris(hydroxymethyl) aminomethane, with the formula $(\text{HOCH}_2)_3\text{CNH}_2$. Tris has the ability to absorb counter ions (+H and -OH) so as to help keep the solution that they are at a stable pH level. When the pH of Tris is set using HCl (hydrochloric acid) the buffer is called Tris HCl.

Tris is extensively used in biochemistry and molecular biology. In biochemistry, Tris is widely used as a component of buffer solutions, such as in TAE and TBE buffer, especially for solutions of nucleic acids. It is a primary amine and thus undergoes the reactions associated with typical amines, e.g. condensations with aldehydes

Application : The useful buffer range for tris (pH 7 ~ 9) coincides with the typical physiological pH of most living organisms. This, and its low cost, make tris one of the most common buffers used in the biology/biochemistry laboratory.

Tris is also used as a primary standard to standardize acid solutions for chemical analysis